# WBJEEM - 2014

## PHYSICS

| Q.No. | ✱ | ⌘ | ❖ | ☼ |
|---|---|---|---|---|
| 01 | B | C | A | C |
| 02 | D | C | D | A |
| 03 | B | D | A | C |
| 04 | C | A | B | C |
| 05 | C | A | A | D |
| 06 | A | D | D | C |
| 07 | B | A | C | B |
| 08 | A | A | D | A |
| 09 | C | D | B | B |
| 10 | D | B | C | A |
| 11 | C | D | A | D |
| 12 | C | B | C | A |
| 13 | B | B | D | A |
| 14 | D | B | A | D |
| 15 | B | D | C | A |
| 16 | B | B | B | A |
| 17 | B | C | A | D |
| 18 | B | B | D | A |
| 19 | A | D | D | C |
| 20 | B | D | C | D |
| 21 | C | D | C | C |
| 22 | A | C | D | D |
| 23 | D | A | C | B |
| 24 | A | C | D | B |
| 25 | A | C | A | A |
| 26 | D | A | A | B |
| 27 | B | B | B | C |
| 28 | C | B | B | D |
| 29 | A | D | B | B |
| 30 | D | A | D | B |
| 31 | D | C | D | B |
| 32 | C | C | C | C |
| 33 | D | A | D | B |
| 34 | D | D | B | A |
| 35 | A | C | A | B |
| 36 | A | A | B | D |
| 37 | C | B | C | C |
| 38 | D | A | D | C |
| 39 | D | C | C | D |
| 40 | D | B | A | D |
| 41 | A | D | A | A |
| 42 | C | A | B | D |
| 43 | A |  | C | C |
| 44 |  | C | C | C |
| 45 | C | D | A | D |
| 6 | B | A | C | D |
| 7 | C | C | C | B |
| 4 | D | A | D | C |
| 49 | D | C | D | B |
| 50 | B | D | B | B |
| 51 | D | B | A | D |
| 52 | C | D | A | A |
| 53 | B | B | B | A |
| 54 | A | B | D | D |
| 55 | A | D | B | C |
| 56 | A, D | B, C, D | A, C | C, D |
| 57 | A, C | A, C | A, C | B, C, D |
| 58 | C, D | C, D | A, D | A, D |
| 59 | A, C | A, C | B, C, D | A, C |
| 60 | B, C, D | A, D | C, D | A, C |

## CHEMISTRY

| Q.No. | ↑ | ← | ↓ | → |
|---|---|---|---|---|
| 01 | A | D | D | A |
| 02 | A | B | C | B |
| 03 | C | C | D | C |
| 04 | C | C | A | A |
| 05 | D | B | B | C |
| 06 | A | A | C | A |
| 07 | A | D | B | B* |
| 08 | C | B | C | C |
| 09 | B | A | C | D |
| 10 | A | C | A | C |
| 11 | D | B | B | B |
| 12 | D | D | A | C |
| 13 | D | A | C | A |
| 14 | C | C | B* | A |
| 15 | C | B* | A | B |
| 16 | B | B | B | D |
| 17 | B |  | D | A |
| 18 | A | C | D | B |
| 19 | A | C | A | A |
| 20 | A | D | D | C |
| 21 | B | C | D | D |
| 22 |  | C | B | C |
| 23 | B | B | C | B |
| 24 | A | A | A | D |
| 25 | C | B | A | B |
| 26 |  | A | A | D |
| 27 | C | A | A | A |
| 28 | B | A | D | D |
| 29 | C | C | D | C |
| 30 | D | B | A | C |
| 31 | B | B | A | B |
| 32 | A | A | B | C |
| 33 | D | D | C | A |
| 34 | C | B | A | B |
| 35 | B | B | B | A |
| 36 | A | D | C | B |
| 37 | D | D | B | B |
| 38 | C | D | B | C |
| 39 | B | C | D | B |
| 40 | D | A | C | A |
| 41 | D | B | C | A |
| 42 | C | A | B | B |
| 43 | B | A | B | D |
| 44 | B | B | B | D |
| 45 | B | B | C | B |
| 46 | B | C | C | A |
| 47 | C | A | D | B |
| 48 | C | A | A | C |
| 49 | D | D | C | C |
| 50 | B | D | B | C |
| 51 | A | C | B | C |
| 52 | C | B | C | A |
| 53 | C | C | B | D |
| 54 | A | B | C | B |
| 55 | B | C | A | B |
| 56 | A, B, D | B, D | B, C, D | B, C |
| 57 | B, C, D, | B, C | A, B, D | A, B, D |
| 58 | B, C | A, B, D | A, B, D | B, D |
| 59 | A, B, D | A, B, D | B, D | B, C, D |
| 60 | B, D | B, C, D | B, C | A, B, D |

*** B and C both option are correct but as single Option B is more appropriate.**

**Code - ✱**

# ANSWERS & HINTS
## *for*
## WBJEEM - 2014
## SUB : PHYSICS

**CATEGORY - I**

**Q.1 to Q.45 carry one mark each, for which only one option is correct. Any wrong answer will lead to deduction of 1/3 mark.**

1. A whistle whose air column is open at both ends has a fundamental frequency of 5100 Hz. If the speed of sound in air is 340 $ms^{-1}$, the length of the whistle, in cm, is

(A) 5/3 (B) 10/3 (C) 5 (D) 20/3

**Ans : (B)**

**Hints :** $f = \frac{v}{2\ell} \Rightarrow \ell = \frac{v}{2f} = \frac{340}{2 \times 5100} = \frac{1}{30}m = \frac{10}{3}cm$

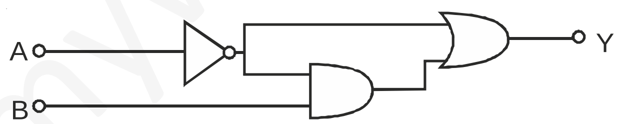
2. One mole of an ideal monoatomic gas is heated at a constant pressure from 0°C to 100°C. Then the change in the internal energy of the gas is (Given R = 8.32 $Jmol^{-1}K^{-1}$)

(A) $0.83 \times 10^3$ J (B) $4.6 \times 10^3$ J (C) $2.08 \times 10^3$ J (D) $1.25 \times 10^3$ J

**Ans : (D)**

**Hints :** $\Delta U = nC_v\Delta T = 1\times\left(\frac{3}{2}R\right)\times 100 = 1\times\frac{3}{2}\times 8.32\times 100 = 1.25 \times 10^3 J$

3. The output Y of the logic circuit given b low is,




(A) $\overline{A}+B$ (B) $\overline{A}$ (C) $\overline{(\overline{A}+B)}.\overline{A}$ (D) $\overline{(\overline{A}+B)}.A$

**Ans : (B)**

**Hints :** $\left(\overline{A}.B\right)+\overline{A} = \overline{A}.(B+1) = \overline{A}.1 = \overline{A}$

4. In which of the following pairs, the two physical quantities have different dimensions?

(A) Planck's constant and angular momentum (B) Impulse and linear momentum

(C) Moment of inertia and moment of a force (D) Energy and torque

**Ans : (C)**

5. A small metal sphere of radius *a* is falling with a velocity $\nu$ through a vertical column of a viscous liquid. If the coefficient of viscosity of the liquid is $\eta$, then the sphere encounters an opposing force of

(A) $6\pi\eta a^2\nu$ (B) $\frac{6\eta\nu}{\pi a}$ (C) $6\pi\eta a\nu$ (D) $\frac{\pi\eta\nu}{6a^3}$

**Ans : (C)**

**Hints :** Stoke's Law

6. A cricket ball thrown across a field is at heights $h_1$ and $h_2$ from the point of projection at times $t_1$ and $t_2$ respectively after the throw. The ball is caught by a fielder at the same height as that of projection. The time of flight of the ball in this journey is

(A) $\dfrac{h_1t_2^2 - h_2t_1^2}{h_1t_2 - h_2t_1}$ (B) $\dfrac{h_1t_1^2 + h_2t_2^2}{h_2t_1 + h_1t_2}$ (C) $\dfrac{h_1t_2^2 + h_2t_1^2}{h_1t_2 + h_2t_1}$ (D) $\dfrac{h_1t_1^2 - h_2t_2^2}{h_1t_1 - h_2t_2}$

**Ans : (A)**

**Hints :** $h_1 = (u\sin\theta)t_1 - \dfrac{1}{2}gt_1^2$ ; $h_2 = (u\sin\theta)t_2 - \dfrac{1}{2}gt_2^2$

$$\Rightarrow \frac{h_1 + \frac{1}{2}gt_1^2}{h_2 + \frac{1}{2}gt_2^2} = \frac{t_1}{t_2} \Rightarrow h_1t_2 - h_2t_1 = \frac{g}{2}\left(t_1t_2^2 - t_1^2t_2\right)$$

$$T = \frac{2u\sin\theta}{g} = \frac{2}{g}\left[\frac{h_1 + \frac{1}{2}gt_1^2}{t_1}\right] = \frac{2}{t_1}\left[\frac{h_1}{g} + \frac{t_1^2}{2}\right] = \frac{h_1}{t_1} \times \left(\frac{t_1t_2^2 - t_1^2t_2}{h_1t_2 - h_2t_1}\right) + t_1 = \frac{h_1t_2^2 - h_2t_1^2}{h_1t_2 - h_2t_1}$$

7. A smooth massless string passes over a smooth fixed pulley. Two masses $m_1$ and $m_2$ ($m_1 > m_2$) are tied at the two ends of the string. The masses are allowed to move under gravity starting from rest. The total external force acting on the two masses is

(A) $(m_1 + m_2)\,g$ (B) $\dfrac{(m_1 - m_2)^2}{m_1 + m_2}g$ (C) $(m_1 - m_2)\,g$ (D) $\dfrac{(m_1 + m_2)^2}{m_1 - m_2}g$

**Ans : (B)**

**Hints :** $a_{cm} = \left(\dfrac{m_1 - m_2}{m_1 + m_2}\right)^2 g$

so, Resultant external force = $(m_1 + m_2)\,a_{cm} = \dfrac{(m_1 - m_2)^2}{(m_1 + m_2)}g$

8. To determine the coefficient of friction between a rough surface and a block, the surface is kept inclined at 45° and the block is released from rest. The block takes a time *t* in moving a distance *d*.The rough surface is then replaced by a smooth surface and the same experiment is repeated. The block now takes a time *t/2* in moving down the same distance *d*. The coefficient of friction is

(A) 3/4 (B) 5/4 (C) 1/2 (D) $1/\sqrt{2}$

**Ans : (A)**

**Hints :** $\mu = \tan\theta\left(1 - \dfrac{1}{n^2}\right) = 1\left[1 - \dfrac{1}{2^2}\right] = \dfrac{3}{4}$

9. A wooden block is floating on water kept in a beaker. 40% of the block is above the water surface. Now the beaker is kept inside a lift that starts going upward with acceleration equal to g/2. The block will then

(A) sink (B) float with 10% above the water surface

(C) float with 40% above the water surface (D) float with 70% above the water surface

**Ans : (C)**

10. An electron in a circular orbit of radius .05 nm performs $10^{16}$ revolutions per second. The magnetic moment due to this rotation of electron is (in $Am^2$)

(A) $2.16 \times 10^{-23}$ (B) $3.21 \times 10^{-22}$ (C) $3.21 \times 10^{-24}$ (D) $1.26 \times 10^{-23}$

**Ans : (D)**

**Hints :** $M = iA = qfA = (1.6 \times 10^{-19})(10^{16})(3.14 \times (0.05\times10^{-9})^2) = 1.26 \times 10^{-23}$

11. A very small circular loop of radius *a* is initially (at t = 0) coplanar and concentric with a much larger fixed circular loop of radius *b*. A constant current *I* flows in the larger loop. The smaller loop is rotated with a constant angular speed $\omega$ about the common diameter. The emf induced in the smaller loop as a function of time *t* is

(A) $\dfrac{\pi a^2 \mu_o I}{2b}\omega\cos(\omega t)$ (B) $\dfrac{\pi a^2 \mu_o I}{2b}\omega\sin(\omega^2 t^2)$

(C) $\dfrac{\pi a^2 \mu_o I}{2b}\omega\sin(\omega t)$ (D) $\dfrac{\pi a^2 \mu_o I}{2b}\omega\sin^2(\omega t)$

**Ans : (C)**

**Hints :** $\varepsilon = NBA\omega \sin\omega t$ N = 1, $B = \dfrac{\mu_0 I}{2b}$, $A = \pi a^2$

$= \dfrac{\mu_0 I}{2b}(\pi a^2)\omega \sin\omega t$

12. A drop of some liquid of volume 0.04 $cm^3$ is placed on the surface of a glass slide. Then another glass slide is placed on it in such a way that the liquid forms a thin layer of area 20 $cm^2$ between the surfaces of the two slides. To separate the slides a force of $16\times10^5$ dyne has to be applied normal to the surfaces. The surface tension of the liquid is (in dyne-$cm^{-1}$)

(A) 60 (B) 70 (C) 80 (D) 90

**Ans : (C)**

**Hints :** Let thickness of layer is t

$V = At,\quad t = \dfrac{V}{A},\quad 2r = \dfrac{V}{A},\ r = \dfrac{V}{2A},\quad \Delta P = \dfrac{T}{r}$

$F = \Delta P \times A = \dfrac{T}{r}\times A = \dfrac{T}{\left(\dfrac{V}{2A}\right)}A,\ F = \dfrac{2TA^2}{V} = $ 80 dyne/cm

13. A proton of mass *m* and charge *q* is moving in a plane with kinetic energy *E.* If there exists a uniform magnetic field *B,* perpendicular to the plane of the motio th portion will move in a circular path of radius

(A) $\dfrac{2Em}{qB}$ (B) $\dfrac{\sqrt{2Em}}{qB}$ (C) $\dfrac{\sqrt{Em}}{2qB}$ (D) $\sqrt{\dfrac{2Eq}{mB}}$

**Ans : (B)**

**Hints :** $r = \dfrac{mv}{qB} = \dfrac{\sqrt{2Em}}{qB}$

14. In which of the following phenomena, the heat waves travel along straight lines with the speed of light ?

(A) thermal conduction (B) forced convection (C) natural convection (D) thermal radiation

**Ans : (D)**

15. An artificial satellite moves in a circular orbit around the earth. Total energy of the satellite is given by *E.* The potential energy of the satellite is

(A) –2E (B) 2E (C) 2E/3 (D) –2E/3

**Ans : (B)**

**Hints :** P.E. = 2(T.E.)

16. A particle moves with constant acceleration along a straight line starting from rest. The percentage increase in its displacement during the 4$^{th}$ second compared to that in the 3$^{rd}$ second is

(A) 33% (B) 40% (C) 66% (D) 77%

**Ans : (B)**

**Hints :** $S_{nth} = u + \frac{1}{2}a(2n-1)$

$S_{3rd} = \frac{5}{2}a\ ,\quad S_{4\ h} = \frac{7}{2}a$

$\frac{S_{4\ h} - S_{3rd}}{S_{3rd}} \times 100 = \frac{a}{\left(\frac{5a}{2}\right)} \times 100 = 40\%$

17. In the circuit shown assume the diode to be ideal. When $V_i$ increases from 2 V to 6 V, the change in the current is (in mA)


(A) zero (B) 20 (C) 80/3 (D) 40

**Ans : (B)**

**Hints :** $I_{initial}$ = 0 , $I_{final}$ = 3/150 = 0.02A

S, change in I = 0.02A = 20 mA

18. In a transistor output characteristics commonly used in common emitter con ig ration, the base current $I_B$, the collector current $I_C$ and the collector-emitter voltage $V_{CE}$ have values of the following orders of magnitude in the active region

(A) $I_B$ and $I_C$ both are in μA and $V_{CE}$ in Volts
(B) $I_B$ is in μA, and $I_C$ is in mA and $V_{CE}$ in Volts
(C) $I_B$ is in mA, and $I_C$ is in μA and $V_{CE}$ in mV
(D) $I_B$ is in mA, nd $I_C$ is in mA and $V_{CE}$ in mV

**Ans : (B)**

19. If *n* denotes a positive integer, *h* the Planck's constant, *q* the charge and *B* the magnetic field, then the quantity $\left(\frac{nh}{2\pi qB}\right)$ has the dimension of

(A) area (B) length (C) speed (D) acceleration

**Ans : (A)**

**Hints :** $\left[\frac{nh}{2\pi qB}\right] = \frac{[mvr]}{[qB]} = \frac{[mvr][v]}{[F]} = \frac{[mv\ \ r]}{\left[\frac{mv^2}{r}\right]} = [r^2]$

20. For the radioactive nuclei that undergo either $\alpha$ or $\beta$ decay, which one of the following cannot occur ?

(A) isobar of original nucl us is produced
(B) isotope of the original nucleus is produced
(C) nuclei with higher atomic number than that of the original nucleus is produced
(D) nuclei with lower atomic number than that of the original nucleus is produced

**Ans : (B)**

21. A car moving with a speed of 72 km-hour$^{-1}$ towards a roadside source that emits sound at a frequency of 850 Hz. The car driver listens to the sound while approaching the source and again while moving away from the source after crossing it. If the velocity of sound is 340 ms$^{-1}$,the difference of the two frequencies, the driver hears is

(A) 50 Hz (B) 85 Hz
(C) 100 Hz (D) 150 Hz

**Ans : (C)**

**Hints :** $\nu_{approach} = \nu\left(\frac{V+Vo}{V}\right) = 850\left(\frac{340+20}{340}\right)$, $\nu_{separation} = 850\left(\frac{340-20}{340}\right)$, $\nu_{approach} - \nu_{separation} =$

$\frac{850}{340} \times 40 = 100$ Hz

22. Same quantity of ice is filled in each of the two metal containers P and Q having the same size, shape and wall thickness but made of different materials. The containers are kept in identical surroundings. The ice in P melts completely in time $t_1$ whereas that in Q takes a time $t_2$. The ratio of thermal conductivities of the materials of P and Q is

(A) $t_2 : t_1$ (B) $t_1 : t_2$ (C) $t_1^2 : t_2^2$ (D) $t_2^2 : t_1^2$

**Ans : (A)**

**Hints :** $\left(KA\frac{dT}{dx}\right)t = mL$, $K \alpha \frac{1}{t}$ So, $\frac{K_1}{K_2} = \frac{t_2}{t_1}$

23. Three capacitors, 3μF, 6μF and 6μF are connected in series to a source of 120V. The potential difference, in volts, across the 3μF capacitor will be

(A) 24 (B) 30 (C) 40 (D) 60

**Ans : (D)**

**Hints :** $Q=CV \Rightarrow V=\frac{Q}{C} \Rightarrow V\alpha\frac{1}{C}$, so, $V = 120\left(\frac{1/3}{1/3 + 1/6 + 1/6}\right) = 60$ volts

24. A galvanometer having internal resistance 10Ω requires 0.01 A for a full scale def ection. To convert this galvanometer to a voltmeter of full-scale deflection at 120V, we need to connect a resistance of

(A) 11990 Ω in series (B) 11990 Ω in parallel (C) 12010 Ω in series (D) 12010 Ω in parallel

**Ans : (A)**

**Hints :** $R = \frac{V}{I_g} - R_g = \frac{120}{0.01} - 10 = 11990\ \Omega$

25. Consider three vectors $\vec{A} = \hat{i} + \hat{j} - 2\hat{k}$, $\vec{B} = \hat{i} - \hat{j} + \hat{k}$ and $\vec{C} = 2\hat{i} - 3\hat{j} + 4\hat{k}$. A vector $\vec{X}$ of the form $\alpha\vec{A} + \beta\vec{B}$ ($\alpha$ and $\beta$ are numbers) is perpendicular to $\vec{C}$. The ratio of $\alpha$ and $\beta$ is

(A) 1:1 (B) 2:1 (C) –1:1 (D) 3:1

**Ans : (A)**

**Hints :** $\left(\alpha\vec{A} + \beta\vec{B}\right)\cdot\vec{C} = 0$, $\Rightarrow 2(\alpha+\beta) - 3(\alpha-\beta)\ \ 4(\beta-2\alpha) = 0$, $\Rightarrow -9\alpha + 9\beta = 0$, $\Rightarrow \alpha:\beta = 1:1$

26. A parallel plate capacitor is charged an then disconnected from the charging battery. If the plates are now moved farther apart by pulling at them by means of insulating handles, then

(A) the energy stored in the capacitor decreases (B) the capacitance of the capacitor increases
(C) the charge on the capacitor decreases (D) the voltage across the capacitor increases

**Ans : (D)**

**Hints :** $C = \frac{\varepsilon_o A}{d}$, $d\uparrow$, $c\downarrow$, Q(Const), $V\uparrow$

27. When a particle executing SHM oscillates with a frequency ν, then the kinetic energy of the particle

(A) changes periodically with a frequency of ν (B) changes periodically with a frequency of 2ν
(C) changes periodically with a frequency of ν/2 (D) remains constant

**Ans : (B)**

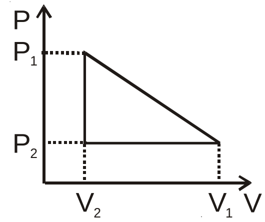
28. The ionization energy of hydrogen is 13.6eV. The energy of the photon released when an electron jumps from the first excited state (n=2) to the ground state of a hydrogen atom is

(A) 3.4 eV (B) 4.53 eV (C) 10.2 eV (D) 13.6 eV

**Ans : (C)**

**Hints :** $13.6\left[\frac{1}{1^2} - \frac{1}{2^2}\right] = 13.6\left(1 - \frac{1}{4}\right) = 13.6 \times \frac{3}{4} = 10.2$ ev

29. One mole of a van der Waals' gas obeying the equation $\left(P+\frac{a}{V^2}\right)(V-b) = RT$ undergoes the quasi-static cyclic process which is shown in the P-V diagram. The net heat absorbed by the gas in this process is




(A) $\frac{1}{2}(P_1-P_2)(V_1-V_2)$  (B) $\frac{1}{2}(P_1+P_2)(V_1-V_2)$

(C) $\frac{1}{2}\left(P_1+\frac{a}{v_1^2}-P_2-\frac{a}{v_2^2}\right)(V_1-V_2)$  (D) $\frac{1}{2}\left(P_1+\frac{a}{v_1^2}+P_2+\frac{a}{v_2^2}\right)(V_1-V_2)$

**Ans : (A)**

**Hints :** For cyclic process, heat absorbed = work done = Area = $\frac{1}{2}(P_1-P_2)(V_1-V_2)$

30. A scientist proposes a new temperature scale in which the ice point is 25 X (X is the new unit of temperature) and the steam point is 305 X. The specific heat capacity of water in this new scale is (in $Jkg^{-1}X^{-1}$)

(A) $4.2\times10^3$  (B) $3.0\times10^3$  (C) $1.2\times10^3$  (D) $1.5\times10^3$

**Ans : (D)**

**Hints :** (305–25)X = 100°C, ⇒ 1°C = 2.8X, Sp. heat capacity of water = $4200\,\frac{J}{Kg\,°C}$, $= 4200\,\frac{J}{Kg\,(2.8X)}$,

$= 1.5\times10^3\ J/(Kg - X)$

31. A metal rod is fixed rigidly at two ends so as to prevent its thermal expansion. If L, $\alpha$ and Y respectively denote the length of the rod, coefficient of linear thermal expansion and Young's modulus of its material, then for an increase in temperature of the rod by $\Delta T$, the longitudinal stress developed in the rod is

(A) inversely proportional to $\alpha$

(B) inversely proportional to Y

(C) directly proportional to $\frac{\Delta T}{Y}$

(D) independent of L

**Ans : (D)**

**Hints :** Strain = $\alpha\Delta T$

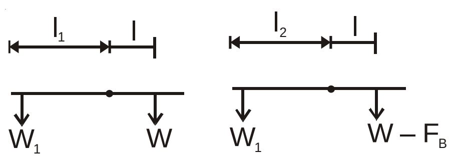
Stress = $Y\alpha\Delta T$

32. A uniform rod is suspended horizontally from its mid-point. A piece of metal whose weight is W is suspended at a distance $l$ from the mid-point. Another weight $W_1$ is suspended on the other side at a distance $l_1$ from the mid-point to bring the rod to a horizontal position. When $W$ is completely immersed in water, $W_1$ needs to be kept at a distance $l_2$ from the mid-point to get the rod back into horizontal position. The specific gravity of the metal piece is

(A) $\frac{W}{W_1}$  (B) $\frac{Wl_1}{Wl - W_1 l_2}$  (C) $\frac{l_1}{l_1 - l_2}$  (D) $\frac{l_1}{l_2}$

**Ans : (C)**

**Hints :**




$\rho$ = specific gravity

$Wl = W_1 l_1$ $\qquad$ $W - F_B = W(1 - 1/\rho)$

$Wl\,(1 - 1/\rho) = W_1 l_2$

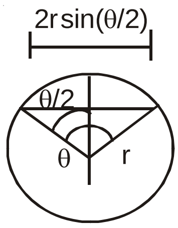
$1 - 1/\rho = \dfrac{l_2}{l_1}$ $\qquad \Rightarrow 1/\rho = 1 - \dfrac{l_2}{l_1} = \dfrac{l_1 - l_2}{l_1}$

$\Rightarrow \rho = \dfrac{l_1}{l_1 - l_2}$

33. A particle is moving uniformly in a circular path of radius r. When it moves through an angular displacement $\theta$, then the magnitude of the corresponding linear displacement will be

(A) $2r\cos\left(\dfrac{\theta}{2}\right)$ (B) $2r\cot\left(\dfrac{\theta}{2}\right)$ (C) $2r\tan\left(\dfrac{\theta}{2}\right)$ (D) $2r\sin\left(\dfrac{\theta}{2}\right)$

**Ans : (D)**

**Hints :** 


34. A luminous object is separated from a screen by distance *d*. A convex len  is placed between the object and the screen such that it forms a distinct image on the screen. The maximum possible focal length of this convex lens is

(A) 4d (B) 2d (C) d/2 (D) d/4

**Ans : (D)**

**Hints :** From lens displacement method

35. The intensity of magnetization of a bar magnet is 5 0 × 0$^4$ $Am^{-1}$. The magnetic length and the area of cross section of the magnet are 12 cm and 1 $cm^2$ respectively The magnitude of magnetic moment of this bar magnet is (in SI unit)

(A) 0.6 (B) 1.3 (C) 1.24 (D) 2.4

**Ans : (A)**

**Hints :** $I = \dfrac{M}{V} \Rightarrow M = IV = 5.0 \times 10^4 \times 12 \times 10^{-6} = 60 \times 10^{-2} = 0.6$

36. An infinite sheet carrying a unif rm surface charge density $\sigma$ lies on the xy-plane. The work done to carry a charge q from the point $\vec{A} = a\left(\hat{i} + 2\hat{j} + 3\hat{k}\right)$ to the point $\vec{B} = a\left(\hat{i} - 2\hat{j} + 6\hat{k}\right)$ (where a is a constant with the dimension of length and $\varepsilon_0$ is the permittivity of free space) is

(A) $\dfrac{3\sigma aq}{2\varepsilon_0}$ (B) $\dfrac{2\sigma aq}{\varepsilon_0}$ (C) $\dfrac{5\sigma aq}{2\varepsilon_0}$ (D) $\dfrac{3\sigma aq}{\varepsilon_0}$

**Ans : (A)**

**Hints :** $\overrightarrow{AB} = a(-4\hat{j} + 3\hat{k})$

$$\text{Workdone} = q\left(\frac{\sigma}{2\varepsilon_0}\right)\hat{k}\cdot a\left(-4\hat{j} + 3\hat{k}\right) = \frac{3q\sigma a}{2\varepsilon_0}$$

37. A uniform solid spherical ball is rolling down a smooth inclined plane from a height h. The velocity attained by the ball when it reaches the bottom of the inclined plane is v. If the ball is now thrown vertically upwards with the same velocity v, the maximum height to which the ball will rise is

(A) 5h/8 (B) 3h/5 (C) 5h/7 (D) 7h/9

**Ans : (C)**

**Hints :** $mgh = \frac{1}{2}mv^2\left(1+\frac{k^2}{R^2}\right)$

$\Rightarrow v = \sqrt{\frac{10gh}{7}}$

For vertical projection,
$v^2 - u^2 = 2gh'$

So, $\frac{10}{7}gh = 2gh' \quad \Rightarrow h' = 5h/7$

38. Two coherent monochromatic beams of intensities I and 4I respectively are superposed. The maximum and minimum intensities in the resulting pattern are

(A) 5I and 3I (B) 9I and 3I (C) 4I and I (D) 9I and I

**Ans : (D)**

**Hints :** $\frac{I_{max}}{I_{man}} = \left(\frac{\sqrt{4I}+\sqrt{I}}{\sqrt{4I}-\sqrt{I}}\right)^2 = \left(\frac{3\sqrt{I}}{\sqrt{I}}\right)^2 = \frac{9}{1}$

39. If the bandgap between valence band and conduction band in a material is 0 eV, then the material is

(A) semiconductor (B) good conductor (C) superconductor (D) insulator

**Ans : (D)**

**Hints :** The band gap of 5 eV corresponds to that of an insulator.

40. Consider a blackbody radiation in a cubical box at absol te temperature T. If the length of each side of the box is doubled and the temperature of the walls of the box and that of the radiation is halved, then the total energy

(A) halves (B) doubles (C) quadruples (D) remains the same

**Ans : (D)**

**Hints :** Assuming temperature of the body and cubica box is same initially i.e. T and finally it becomes T/2. Because temperature of body and surrounding remains same Hence no net loss of radiation occur through the body. Thus total energy remains constant.

41. Four cells, each of emf E and intern l resistance r, are connected in series across an external resistance R. By mistake one of the cells is connected in reverse. Then the current in the external circuit is

(A) $\frac{2E}{4r+R}$ (B) $\frac{3E}{4r+R}$ (C) $\frac{3E}{3r+R}$ (D) $\frac{2E}{3r+R}$

**Ans : (A)**

**Hints :** $i = \frac{3E-E}{4r+R} = \frac{2E}{4r+R}$

42. The energy of gamma $(\gamma)$ ray photon is $E_\gamma$ and that of an X-ray photon is $E_x$. If the visible light photon has an energy of $E_v$, then we can say that

(A) $E_x > E_\gamma > E_v$ (B) $E_\gamma > E_v > E_x$ (C) $E_\gamma > E_x > E_v$ (D) $E_x > E_v > E_\gamma$

**Ans : (C )**

43. The intermediate image formed by the objective of a compound microscope is

(A) real, inverted and magnified (B) real, erect and magnified
(C) virtual, erect and magnified (D) virtual, inverted and magnified

**Ans : ( A)**

44. The displacement of a particle in a periodic motion is given by $y = 4\cos^2\left(\frac{t}{2}\right)\sin(1000t)$. This displacement may be considered as the result of superposition of n independent harmonic oscillations, Here n is

(A) 1 (B) 2 (C) 3 (D) 4

**Ans : (C)**

**Hints :** $y = 4\cos^2\left(\frac{t}{2}\right)\sin(1000t) = 2(1+\cos t)\sin(1000t) = 2\sin 1000t + 2\cos t \cdot \sin 1000t$

$= 2\sin 1000\,t + \sin(1001\,t) + \sin(999\,t)$

45. Consider two concentric spherical metal shells of radii $r_1$ and $r_2$ ($r_2 > r_1$). If the outer shell has a charge q and the inner one is grounded, the charge on the inner shell is

(A) $\frac{-r_2}{r_1}q$ (B) zero (C) $\frac{-r_1}{r_2}q$ (D) –q

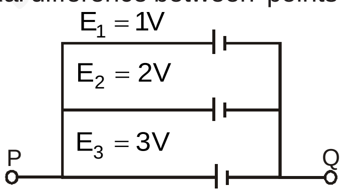
**Ans : (C)**

**Hints :** $\frac{k\,q'}{r_1} + \frac{k\,q}{r_2} = 0 \Rightarrow q' = -\left(\frac{r_1}{r_2}\right)q$

**CATEGORY - II**

**Q.46 to Q.55 carry two marks each, for which only one option is correct. Any wrong answer will lead to deduction of 2/3 mark**

46. A circuit consists of three batteries of emf $E_1$ = 1 V, $E_2$ = 2 V and $E_3$ = 3 V and internal resistances 1 Ω , 2 Ω and 1 Ω respectively which are connected in parallel as shown in the figure. The potential difference between points P and Q is




(A) 1.0 V (B) 2.0 V

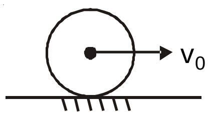
(C) 2.2 V (D) 3.0 V

**Ans : (B)**

**Hints :** $E_{eff} = \dfrac{\frac{1}{1}+\frac{2}{2}+\frac{3}{1}}{\left(\frac{1}{1}+\frac{1}{2}+\frac{1}{1}\right)} = \frac{5}{5}\times 2 = 2$ volt

P.D between two point P and Q = 2 volt

47. A solid uniform sphere resting on a rough horizontal plane is given a horizontal impulse directed through its center so that it starts sliding with an initial vel city $v_0$. When it finally starts rolling without slipping the speed of its center is




(A) $\frac{2}{7}v_0$ (B) $\frac{3}{7}v_0$

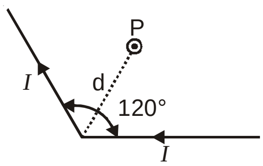
(C) $\frac{5}{7}v_0$ (D) $\frac{6}{7}v_0$

**Ans : (C)**

**Hints :** Angular momentum will remain conserved along point of contact

$mv_0R = mvR + \frac{2}{5}mR^2\left(\frac{v}{R}\right) \Rightarrow v = \frac{5v_0}{7}$

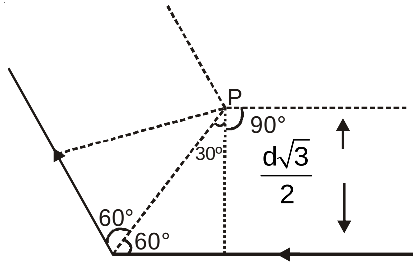
48. A long conducting wire carrying a current I is bent at 120°(see figure). The magnetic field B at a point P on the right bisector of bending angle at a distance d from the bend is ($\mu_0$ is the permeability of free space)




(A) $\frac{3\mu_0 I}{2\pi d}$ (B) $\frac{\mu_0 I}{2\pi d}$

(C) $\frac{\mu_0 I}{\sqrt{3}\pi d}$ (D) $\frac{\sqrt{3}\,\mu_0 I}{2\pi d}$

**Ans : (D)**




$$B_{net} = 2\left[\frac{\mu_0}{4\pi}\times\frac{i}{\left(\frac{d\sqrt{3}}{2}\right)}\times[1+\sin 30°]\right] = 2\left[\frac{\mu_0}{4\pi}\times\frac{2i}{d\sqrt{3}}\times\frac{3}{2}\right] = \frac{\sqrt{3}\mu_0 I}{2\pi d}$$

49. An object is placed 30 cm away from a convex lens of focal length 10 cm and a sharp image is formed on a screen. Now a concave lens is placed in contact with the convex lens. The screen now has to be moved by 45 cm to get a sharp image again. The magnitude of focal length of the concave len is (in cm)

(A) 72 (B) 60 (C) 36 (D) 20

**Ans : (D)**

**Hints :** $\frac{1}{f}=\frac{1}{v}-\frac{1}{u}$ , $\frac{1}{10}=\frac{1}{v}+\frac{1}{30}$ , v = 15 cm. When concave lens is placed v′ = (45 + 15) = 60 cm

$\frac{1}{f}=\frac{1}{v'}-\frac{1}{u}$ (f = focal length of combination), $\frac{1}{f}=\frac{1}{60}+\frac{1}{30}$ = $\boxed{f = 20\text{ m}}$

$\frac{1}{f}=\frac{1}{f_1}+\frac{1}{f_2}$, $\frac{1}{20}=\frac{1}{10}+\frac{1}{f_2}$, $\frac{1}{20}-\frac{1}{10}=\frac{1}{f_2}$ $\boxed{f_2 = -20\text{ m}}$

50. A 10 watt electric heater is used to heat a container filled with 0.5 kg of water. It is found that the temperature of water and the container rises by 3° K in 15 minutes. The contain r is then emptied, dried and filled with 2 kg of oil. The same heater now raises the temperature of container-oil system by 2°K in 20 minutes. Assuming that there is no heat loss in the process and the specific heat of water as 4200 $Jkg^{-1}K^{-}$ , the specific heat of oil in the same unit is equal to

(A) $1.50\times10^3$ (B) $2.55\times10^3$ (C) $3.00\times10^3$ (D) $5.10\times10^3$

**Ans : (B)**

**Hints :** $\left(\frac{1}{2}\times4200\times3\right)+(m_c\times c_c\times3)=10\times15\times60-----(1)$

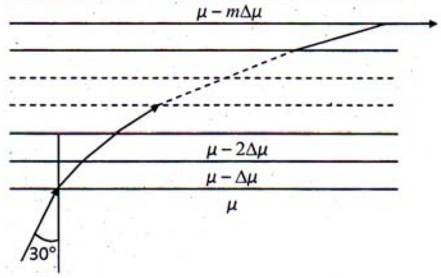
$(m_c\times c_c)=900$. In case of oil. $(2\times c_0\times2)+(m_c\times c_c\times2)=(10\times20\times60)$, $4C_0$ + (900×2)=12000

$(C_0) = 2.55\times10^3\ J\ kg^{-1}k^{-1}$

$C_c$ = Sp. heat capacity of container

$C_0$ = Sp. heat capcity of o l

51. A glass slab consists of thin u iform layers of progressively decreasing refractive indices RI (see figure) such that the RI of any layer is μ-mΔμ. Here μ and Δμ denote the RI of 0th layer and the difference in RI between any two consecutive layers, respectively. The integer m = 0, 1, 2, 3..... denotes the numbers of the successive layers. A ray of light from the 0th layer enters the 1st layer at an angle of incidence of 30°. After undergoing the mth refraction, the ray emerges parallel to the interface. If μ = 1.5 and Δμ = 0.015, the value of m is




(A) 20 (B) 30 (C) 40 (D) 50

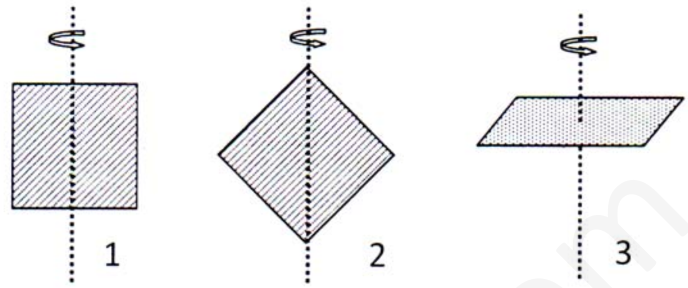
**Ans : (D)**

**Hints :** By Snell's law, $\mu$ sin i = constant,1.5 sin 30° = ($\mu$ – m$\Delta\mu$) sin 90°, $\frac{3}{2}\times\frac{1}{2}=(1.5-m\times0.15)\times1$, $\therefore$ m = 50

52. The de-Broglie wavelength of an electron is the same as that of a 50 keV X-ray photon. The ratio of the energy of the photon to the kinetic energy of the electron is (the energy equivalent of electron mass is 0.5 MeV)
(A) 1 : 50 (B) 1 : 20 (C) 20 : 1 (D) 50 : 1
**Ans : (C)**

**Hints :** $\lambda=\frac{h}{\sqrt{2mK}}$, $K_{electron}=\frac{h^2}{(\lambda^2\times2m)}$, $E_{photon}=\frac{hC}{\lambda}$, $\frac{E_{photon}}{K_{electron}}=\left[\frac{hc}{\lambda}.\frac{\lambda^2\times2m}{h^2}\right]=\frac{2mC^2}{\left(hC/\lambda\right)}=\frac{2\times5\times10^5}{(50\times10^3)}=\frac{20}{1}$

53. Three identical square plates rotate about the axes shown in the figure in such a way that their kinetic energies are equal. Each of the rotation axes passes through the centre of the square. Then the ratio of angular speeds $\omega_1:\omega_2:\omega_3$ is




(A) 1 : 1 : 1 (B) $\sqrt{2}:\sqrt{2}:1$ (C) 1 : $\sqrt{2}$ : 1 (D) $1:2:\sqrt{2}$
**Ans : (B)**

**Hints :** K = $\frac{1}{2}I\omega^2$, $\omega\propto\frac{1}{\sqrt{I}}$, $\omega_1:\omega_2:\omega_3$ = 1 : 1 : $\frac{1}{\sqrt{2}}=\sqrt{2}$ $\sqrt{2}:1$

54. To determine the composition of a bimetallic alloy, a sample is first weighed in air and then in water. These weights are found to be $w_1$ and $w_2$ respectively. If the densities of t e two constituent metals are $\rho_1$ and $\rho_2$ respectively, then the weight of the first metal in the sample is (whe e $\rho_w$ is the density of water)

(A) $\frac{\rho_1}{\rho_w(\rho_2-\rho_1)}[w_1(\rho_2-\rho_w)-w_2\rho_2]$ (B) $\frac{\rho_1}{\rho_w(\rho_2+\rho_1)}[w_1(\rho_2-\rho_w)+w_2\rho_2]$

(C) $\frac{\rho_1}{\rho_w(\rho_2-\rho_1)}[w_1(\rho_2+\rho_w)-w_2\rho_1]$ (D) $\frac{\rho_1}{\rho_w(\rho_2-\rho_1)}[w_1(\rho_1-\rho_w)-w_2\rho_1]$
**Ans : (A)**

**Hints :** ($w_1$–$w_2$) = v $\rho_w$g, ($w_1$–$w_2$) = ($v_1$ + $v_2$) $\rho_w$g, ($w_1$–$w_2$) = $\left[\frac{x}{\rho_1}+\frac{(w_1-x)}{\rho_2}\right]\rho_w$

(x - weight of the first metal) $x=\frac{\rho_1}{\rho_w(\rho_2-\rho_1)}[w_1(\rho_2-\rho_w)-w_2\rho_2]$

55. Sound waves are passing through two routes-one in straight path and the other along a semicircular path of radius r and are again combined into one pipe and superposed as shown in the figure. If the velocity of sound waves in the pipe is v, then frequencies of resultant waves of maximum amplitude will be integral multiples of

(A) $\frac{v}{r(\pi-2)}$ (B) $\frac{v}{r(\pi-1)}$ (C) $\frac{2v}{r(\pi-1)}$ (D) $\frac{v}{r(\pi+1)}$

**Hints :**

Path difference = $(\pi r - 2r) = (\pi - 2)r = n\lambda$

f- frequency. $v = f \times \lambda$, $\frac{v}{\lambda} = f \Rightarrow \left[\frac{v}{(\pi - 2)r}\right]n = f$

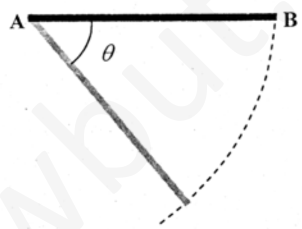
**CATEGORY - III**

**Q.56 to Q.60 carry two marks each, for which one or more than one options may be correct. Marking of correct options will lead to a maximum mark of two on pro rata basis. There will be no negative marking for these questions. However, any marking of wrong option will lead to award of zero mark against the respective question – irrespective of the number of correct options marked.**

56. Find the correct statement(s) about photoelectric effect
   (A) There is no significant time delay between the absorption of a suitable radiation and the emission of electrons
   (B) Einstein analysis gives a threshold frequency above which no electron can be emitted
   (C) The maximum kinetic energy of the emitted photoelectrons is proportional to the frequency of incident radiation
   (D) The maximum kinetic energy of electrons does not depend on the intensity of radiation

   **Ans : (A & D)**

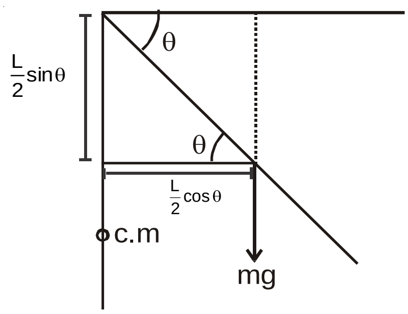
57. A thin rod AB is held horizontally so that it can freely rotate in a vertical plane about the end A as shown in the figure. The potential energy of the rod when it hangs vertically is taken to be zero. The end B of the rod is released from rest from a horizontal position. At the instant the rod makes an angle $\theta$ with the horizontal.




   (A) the speed of end B is proportiona to $\sqrt{\sin\theta}$
   (B) the potential energy is pr portional o $(1–\cos\theta)$
   (C) the angular acceleration is proportional to $\cos\theta$
   (D) the torque about A rem ins the same as its initial value

   **Ans : (A,C)**

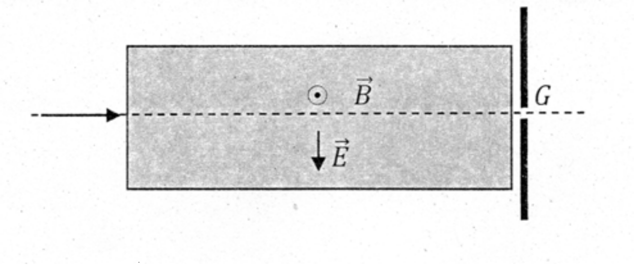
   **Hints :**




Loss in Potential Energy = gain in Kinetic Energy, $mg\frac{L}{2}\sin\theta = \frac{1}{2}I\omega^2$, $\omega \propto \sqrt{\sin\theta}$, $v \propto \sqrt{\sin\theta}$

$U = mgh = mg\,\frac{L}{2}(1 - \sin\theta)$ $\because \tau = I\,\alpha \Rightarrow mg \times \frac{L}{2}\cos\theta = \frac{ml^2}{3} \times \alpha$ . $\alpha \propto \cos\theta$

58. A stream of electrons and protons are directed towards a narrow slit in a screen (see figure). The intervening region has a uniform electric field $\vec{E}$ (vertically downwards) and a uniform magnetic field $\vec{B}$ (out of the plane of the figure) as shown. Then




(A) electrons and protons with speed $\frac{|\vec{E}|}{|\vec{B}|}$ will pass through the slit

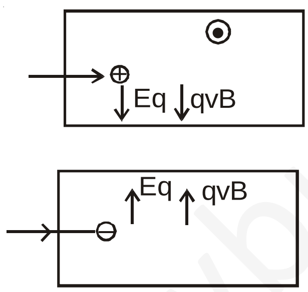
(B) protons with speed $\frac{|\vec{E}|}{|\vec{B}|}$ will pass through the slit, electrons of the same speed will not

(C) neither electrons nor protons will go through the slit irrespective of their speed

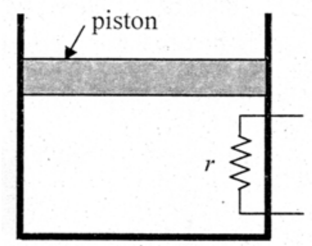
(D) electrons will always be deflected upwards irrespective of their speed

**Ans : (C,D)**

**Hints :**




59. A heating element of resistance r is fitted inside an adiabatic cylinder which carries a frictionless piston of mass m and cross-section A as shown in diagram. The cylinder contains one mole of an ideal diatomic gas. The current flows through the element such tha the temp rature rises with time t as $\Delta T = \alpha t + \frac{1}{2}\beta t^2$ ($\alpha$ and $\beta$ are constants), while pressure remains constan The atmospheric pressure above the piston is $P_0$. Then




(A) the rate of increase in internal energy is $\frac{5}{2}R(\alpha + \beta t)$

(B) the current flowing in the element is $\sqrt{\frac{5}{2r}R(\alpha+\beta t)}$

(C) the piston moves upwards with constant acceleration

(D) the piston moves upwards with constant speed

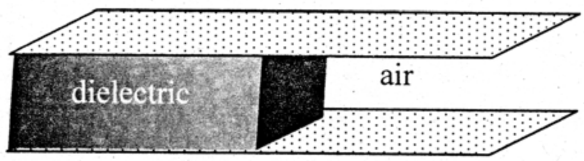
**Ans : (A & C)**

**Hints :** Internal energy $U = \frac{nfRT}{2}$, $U = \frac{5R}{2}\left[\alpha t + \frac{1}{2}\beta t^2\right]$, $\frac{dU}{dt} = \frac{5R}{2}[\alpha + \beta t]$, $dQ = nC_P dT$, $\frac{dQ}{dt} = nC_P \times \frac{dT}{dt}$,

$i^2 r = \frac{7}{2}R \times [\alpha + \beta t]$, $i = \sqrt{\frac{7}{2r}R(\alpha + \beta t)}$, $PV = nRT$, $V = \frac{nRT}{P}$, $V = \frac{nR}{P}\left[\alpha t + \frac{1}{2}\beta t^2\right]$,

$x = \frac{nR}{PA}\left[\alpha t + \frac{1}{2}\beta t^2\right]$, $v = \frac{nR}{PA}[\alpha + \beta t]$, acceleration $= \frac{nR}{PA} \times \beta$

60. Half of the space between the plates of a parallel-plate capacitor is filled with a dielectric material of dielectric constant K. The remaining half contains air as shown in the figure. The capacitor is now given a charge Q. Then




(A) electric field in the dielectric-filled region is higher than that in the air-filled regi n
(B) on the two halves of the bottom plate the charge densities are unequal
(C) charge on the half of the top plate above the air-filled part is $\frac{Q}{K+}$
(D) capacitance of the capacitor shown above is $(1+K)\frac{C_0}{2}$, where $C_0$ is the capacitance of the same capacitor with the dielectric removed

**Ans : (B, C, D)**

**Hints :** $C_1 = \frac{K \epsilon_0 A}{2d}$, $C_2 = \frac{\epsilon_0 A}{2d}$, $C_{eq} = \frac{\epsilon .A}{2d}(K+1) = \frac{C_0}{2}(K+1)$, $\frac{Q_1}{Q_2} = \frac{C_1}{C_2} = \frac{K}{1} \Rightarrow \frac{\sigma_1}{\sigma_2} = \frac{K}{1}$,

$Q_1 = \frac{KQ}{K+1}$ and $Q_2 = \frac{Q}{K+1}$, $E = \frac{\sigma}{\epsilon_0 K}$, $\frac{E_1}{E_2}$ $\frac{\sigma}{\sigma_2} \times \frac{K_2}{K_1} = \frac{Q_1}{Q_2} \times \frac{K_2}{K_1} = \frac{K}{1} \times \frac{1}{K} = 1:1$

❑ ❑ ❑

**Code-↑**

# ANSWERS & HINTS
# *for*
# WBJEEM - 2014
# SUB : CHEMISTRY

**CATEGORY - I**

**Q.1 to Q.45 carry one mark each, for which only one option is correct. Any wrong answer will lead to deduction of 1/3 mark.**

1. During the emission of a positron from a nucleus, the mass number of the daughter element remains the same but the atomic number

(A) is decreased by 1 unit (B) is decreased by 2 units
(C) is increased by 1 unit (D) remains unchanged

**Ans : (A)**

**Hints :** ${}^{A}_{Z}X \rightarrow {}^{A}_{Z-1}Y + {}^{0}_{+1}e$

Atomic number is decreased by 1

2. Four gases P, Q, R and S have almost same values of ‘b but their ‘a’ values (a, b are van der Waals constants) are in the order Q < R < S < P. At a particular temperature among the four gases the most easily liquefiable one is

(A) P (B) Q (C) R (D) S

**Ans : (A)**

**Hints :** More the value of ‘a’ for the gas, more i the intermolecular forces of attraction. Thus the gas can be easily liquefied.

3. $\beta$ emission is always accompanied by

(A) formation of antineutrino and $\alpha$ particle (B) emission of $\alpha$ particle and $\gamma$-ray
(C) formation of antineutrino and $\gamma$-ray (D) formation of antineutrino and positron

**Ans : (C)**

4. The values of $\Delta H$ and $\Delta S$ of a certain reaction are – 400 kJ mol$^{-1}$ and –20 kJ mol$^{-1}$K$^{-1}$ respectively. The temperature below which the reaction is spontaneous is

(A) 100°K (B) 20°C (C) 20°K (D) 120°C

**Ans : (C)**

**Hints :** The reaction is spontaneous when $\Delta G$ is -ve

$\Delta G < 0$

$\Delta H - T\,\Delta S < 0$

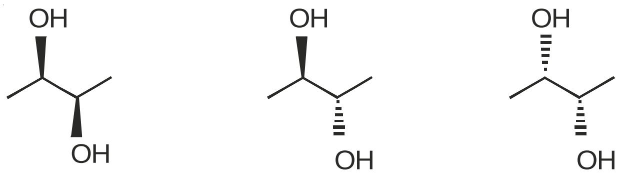
$-400 - (T)(-20) < 0$

$-400 + 20T < 0$

$20T < 400$

$T < \frac{400}{20}$ ; $T < 20K$

5. The correct statement regarding the following compounds is



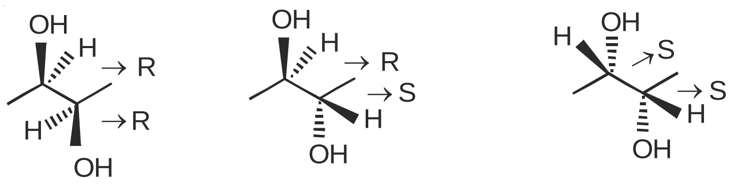
I II III

(A) all three compounds are chiral (B) only I and II are chiral

(C) I and III are diastereomers (D) only I and III are chiral

**Ans : (D)**

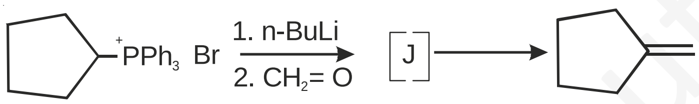
**Hints :**

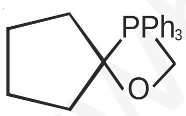
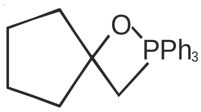
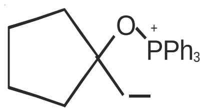


I II III

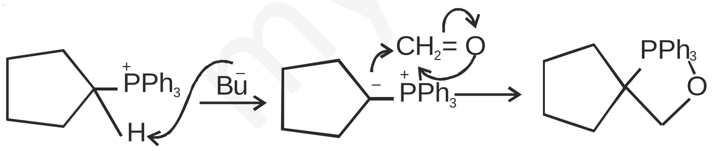
- I and III are enantiomers
- II has plane of symmetry hence achiral

6. The intermediate J in the following Wittig reaction is



(A)  (B)  (C)  (D) 

**Ans : (A)**

**Hints :** 

7. Among the following compounds, the one(s) that gives (give)effervescence with aqueous $NaHCO_3$ solution is (are)

| $(CH_3CO)_2O$ | $CH_3COOH$ | PhOH | $CH_3COCHO$ |
|---|---|---|---|
| **I** | **II** | **III** | **IV** |

(A) I and II (B) I and III (C) only II (D) I and IV

**Ans : (A)**

**Hints :** $CH_3COOH + NaHCO_3 \rightarrow CH_3COONa + CO_2 + H_2O$

$$CH_3-\overset{\overset{\displaystyle O}{\|}}{C}-O-\overset{\overset{\displaystyle O}{\|}}{C}-CH_3 + H_2O \rightarrow 2CH_3COOH$$

$$CH_3COOH + NaHCO_3 \rightarrow CH_3COONa + CO_2 + H_2O$$

8. The system that contains the maximum number of atoms is

(A) 4.25 g of $NH_3$ (B) 8 g of $O_2$ (C) 2 g of $H_2$ (D) 4 g of He

**Ans : (C)**

**Hints :** a) $4.25g\ NH_3 = \left(\frac{4.25}{17}\right)N_A \times 4 = N_A$ atoms

b) $8\ g\ O_2 = \left(\frac{8}{32}\right)N_A \times 2 = \frac{N_A}{2}$ atoms

c) $2\ g\ H_2 = \left(\frac{2}{9}\right)N_A \times 2 = 2N_A$ atoms

d) $4\ g\ He = \left(\frac{4}{4}\right)N_A = N_A$atoms

9. Metal ion responsible for the Minamata disease is

(A) $Co^{2+}$ (B) $Hg^{2+}$ (C) $Cu^{2+}$ (D) $Zn^{2+}$

**Ans : (B)**

**Hints :** $Hg^{2+}$ causes Minamata diseases

10. Among the following observations, the correct one that differentiates between $SO_3^{2-}$ and $SO_4^{2-}$ is

(A) Both form precipitate with $BaCl_2$, $SO_3^{2-}$ dissolves in HCl but $SO_4^{2-}$ does not

(B) $SO_3^{2-}$ forms precipitate with $BaCl_2$, $SO_4^{2-}$ does not

(C) $SO_4^{2-}$ forms precipitate with $BaCl_2$, $SO_3^{2-}$ does not

(D) Both form precipitate with $BaCl_2$, $SO_4^{2-}$ dissolves in HCl but $SO_3^{2-}$ does not

**Ans : (A)**

**Hints :** $BaCl_2 + SO_4^{2-} \rightarrow BaSO_4\downarrow + 2Cl^-$

$BaCl_2 + SO_3^{2-} \rightarrow BaSO_3\downarrow + 2Cl^-$

But $BaSO_3$ dissolves in HCl as $BaSO_3 + 2HCl \rightarrow BaCl_2 + SO_2\uparrow + H_2O$

11. The pH of $10^{-4}$ M KOH solution will be

(A) 4 (B) 11 (C) 10.5 (D) 10

**Ans : (D)**

**Hints :** $[OH^-] = 10^{-4}$ M $\Rightarrow$ pOH = 4

pH + pOH = 14, $\therefore$ pH = 14 - 4 = 10

12. The reagents to carry out the following conversion are

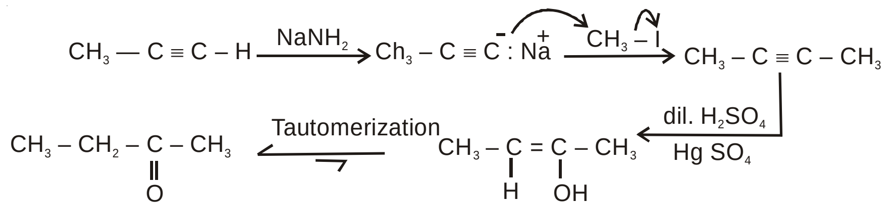

(A) $HgSO_4$/dil $H_2SO_4$ (B) $BH_3$;$H_2O_2$/NaOH

(C) $OsO_4$; $HIO_4$ (D) $NaNH_2/CH_3I$; $HgSO_4$/dil $H_2SO_4$

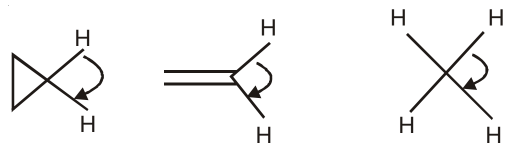
**Ans : (D)**

**Hints :** 

or $CH_3 — C \equiv C – H$




13. The correct order of decreasing H-C-H angle in the following molecules is

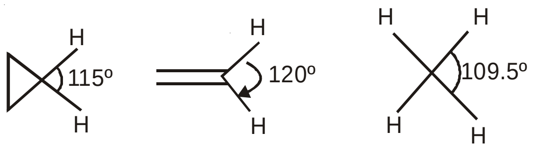



I II III

(A) I > II > III (B) II > I > III (C) III > II > I (D) I > III > II

**Ans : (B)**

**Hints :** II > I > III




I II III

14. $_{98}Cf^{246}$ was formed along with a neutron when an unknown radioactive substance was bombarded with $_6C^{12}$. The unknown substance was

(A) $_{91}Pa^{234}$ (B) $_{90}Th^{234}$ (C) $_{92}U^{235}$ (D) $_{92}U^{238}$

**Ans : (C)**

**Hints :** $_zX^A + {}_6C^{12} \rightarrow {}_{98}Cf^{246} + {}_0n^1$

$z + 6 = 98$ $A + 12 = 246 + 1$

$\Rightarrow z = 92$ or, $A = 247 – 12$

$= 235$

$\therefore$ The element is $_{92}U^{235}$

15. The rate of a certain reaction is given by, rate = k $[H^+]^n$. The rate increases 100 times when the pH changes from 3 to 1. The order (n) of the reaction is

(A) 2 (B) 0 (C) 1 (D) 1.5

**Ans : (C)**

**Hints :** Rate $r = k[H^+]^n$

New rate, $r' = 100\ r$

pH changes from 3 to 1

i.e. $[H^+] = 10^{-3}M$ changes to $[H^+]' = 10^{-1}M$

i.e. conc. increases 100 times $\dfrac{[H^+]'}{[H^+]} = \dfrac{10^{-1}}{10^{-3}} = 100$

$\dfrac{r'}{r} = \left(\dfrac{[H^+]'}{[H^+]}\right)^n$ or, $100 = (100)^n$

or, $n = 1$

16. $({}_{32}Ge^{76}, {}_{34}Se^{76})$ and $({}_{14}Si^{30}, {}_{16}S^{32})$ are examples of

(A) isotopes and isobars
(B) isobars and isotones
(C) isotones and isotopes
(D) isobars and isotopes

**Ans : (B)**

**Hints :** $\left({}_{32}Ge^{76}, {}_{34}Se^{76}\right)$ Same atomic mass = isobars

$\left({}_{14}Si^{30}, {}_{16}Se^{32}\right)$

A – Z = 30 – 14 = 16 Same no. of neutrons = isotones

and 32 – 16 = 16

17. The enthalpy of vaporization of a certain liquid at its boiling point of 35ºC is 24.64 kJ $mol^{-1}$. The value of change in entropy for the process is

(A) 704 J $K^{-1}mol^{-1}$
(B) 80 J $K^{-1}mol^{-1}$
(C) 24.64 J $K^{-1}mol^{-1}$
(D) 7.04 J $K^{-1}mol^{-1}$

**Ans : (B)**

**Hints :** $\Delta S = \frac{q_{rev}}{T}$

At constant pressure, $q_{rev} = \Delta H_{transformation}$

$\Delta S_{vap} = \frac{\Delta H_{vap}}{T_b}$ ; $T_b$ = boiling point, $\Delta H_{vap}$ = Enthalpy of vapourizati n

$= \frac{24.64 \times 10^3 \, Jmol^{-1}}{308 \quad K}$ = 80 $JK^{-1}mol^{-1}$

18. Given that :

$C + O_2 \rightarrow CO_2$ ; $\Delta H^o = -x$ kJ

$2CO + O_2 \rightarrow 2CO_2$ ; $\Delta H^o = -y$ kJ

The heat of formation of carbon monoxide will be

(A) $\frac{y-2x}{2}$
(B) y + 2x
(C) 2x – y
(D) $\frac{2x-y}{2}$

**Ans : (A)**

**Hints :** i) $C + O_2 \rightarrow CO_2$ ; $\Delta H^o = -x$ kJ

ii) $2CO + O_2 \rightarrow 2CO_2$ ; $\Delta H = -y$ kJ

Eq (i) × 2

$2C + 2O_2 \rightarrow 2CO_2$ $\Delta H^o = -2x$ kJ

Writing eq. (ii) in rever e order

$2CO_2 \rightarrow 2CO + O_2$, $\Delta H^o = y$ kJ

adding, $2C + O_2 \rightarrow 2CO$, $\Delta H = (y - 2x)$ kJ

For 2 mol CO, $\Delta H = (y - 2x)$ kJ

$\therefore$ For 1 mol CO, $\Delta H_f = \left(\frac{y-2x}{2}\right)$ kJ

$\therefore$ Enthalpy of formation, $\Delta H^o_f = \frac{y-2x}{2}$

19. Commercial sample of $H_2O_2$ is labeled as 10V. Its % strength is nearly

(A) 3
(B) 6
(C) 9
(D) 12

**Ans : (A)**

**Hints :** 10 volume $H_2O_2$ means

1 mL $H_2O_2$ solution produces 10 mL $O_2$ at STP

$2H_2O_2 \longrightarrow 2H_2O + O_2$

2 mol — 1 mol

2 × 34 g — 22.4 L at STP

68 g

22400 mL $O_2$ at STP is produced from 68 g. $H_2O_2$

$\therefore$ 10 mL $O_2$ is produced from $\dfrac{68 \times 10}{22400}$ g = 0.03036 g $H_2O_2$

$\therefore$ 1 mL $H_2O_2$ solution contains 0.03 g $H_2O_2$ (approx.)

$\therefore$ 100 mL $H_2O_2$ solution contains 0.03 × 100

= 3 g $H_2O_2$ (approx.)

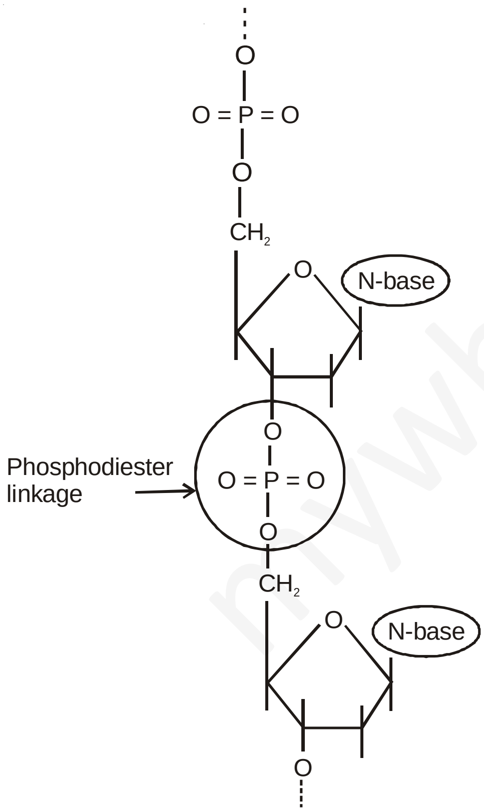
20. In DNA, the consecutive deoxynucleotides are connected via

(A) phospho diester linkage (B) phospho monoester linkage

(C) phospho triester linkage (D) amide linkage

**Ans : (A)**

**Hints :**

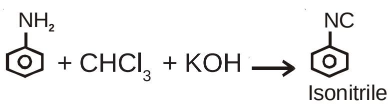



21. The reaction of aniline with chloroform under alkaline conditions leads to the formation of

(A) Phenyl cyanide (B) Phenyl isonitrile (C) Phenyl cyanate (D) Phenyl isocyanate

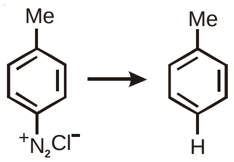
**Ans : (B)**

**Hints :**



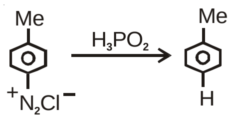

This is the carbylamine reaction

22. The reagent with which the following reaction is best accomplished is



(A) $H_3PO_2$ (B) $H_3PO_3$ (C) $H_3PO_4$ (D) $NaHSO_3$

**Ans : (A)**

**Hints :** 

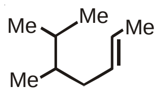
23. At a certain temperature the time required for the complete diffusion of 200 mL of $H_2$ gas is 30 minutes. The time required for the complete diffusion of 50 mL of $O_2$ gas at the same temperature will be

(A) 60 minutes (B) 30 minutes (C) 45 minutes (D) 15 minutes

**Ans : (B)**

**Hints :** $\frac{r_{H_2}}{r_{O_2}} = \frac{\sqrt{M_{O_2}}}{\sqrt{M_{H_2}}} = \frac{V_{H_2}/t_{H_2}}{V_{O_2}/t_{O_2}}$, $\sqrt{\frac{32}{2}} = \frac{200}{30} \times \frac{t_{O_2}}{50} =$ or $4 = \frac{4}{30} \times t_{O_2}$, $\therefore t_{O_2} = 30$ min

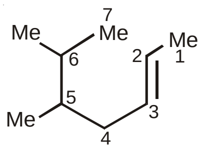
24. The IUPAC name of the following molecule is



(A) 5,6-Dimethyl hept-2-ene (B) 2,3-Dimethyl hept-5-ene

(C) 5,6-Dimethyl hept-3-ene (D) 5-I opropyl hex-2-ene

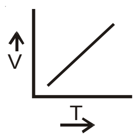
**Ans : (A)**

**Hints :**  5,6 - Dimethyl hept-2 en

25. For one mole of an ideal gas the slope of V vs T curve at constant pressure of 2 atm is X lit mol$^{-1}$K$^{-1}$. The value of the ideal universal gas constant 'R' in term of X is

(A) X lit atm mol$^{-1}$K$^{-1}$ (B) X/2 lit a m mol$^{-1}$K$^{-1}$ (C) 2X lit atm mol$^{-1}$K$^{-1}$ (D) 2X atm lit$^{-1}$mol$^{-1}$K$^{-1}$

**Ans : (C)**

**Hints :** 

$PV = RT$, $V = \frac{R}{P} \times T$, $m = \frac{R}{P} = X$, or $R = X.P$, $= 2X$ L.atm mol$^{-1}$K$^{-1}$

('m' is the slope)

26. An atomic nucleus having low n/p ratio tries to find stability by

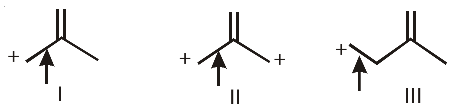
(A) the emission of an $\alpha$ particle (B) the emission of a positron

(C) capturing an orbital electron (K-electron capture) (D) emission of a $\beta$ particle

**Ans : (B)**

**Hints :** B and C both option are correct but as single option, B is more appropriate.

27. The correct order of decreasing length of the bond as indicated by the arrow in the following structure is

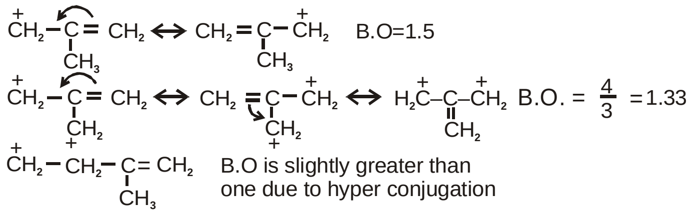


I II III

(A) I>II>III (B) II>I>III (C) III>II>I (D) I>III>II

**Ans : (C)**

**Hints :** I: II: III:



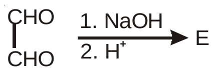

28. If $Cl_2$ is passed through hot aqueous NaOH, the products formed have Cl in different oxidation states. These are indicated as

(A) –1 and +1 (B) –1 and +5 (C) +1 and +5 (D) –1 and +3

**Ans : (B)**

**Hints :** Reaction : $3Cl_2 + 6NaOH \text{ (hot \& conc)} \rightarrow 5\,Na\overset{-1}{Cl} + Na\overset{+5}{Cl}O_3 + 3H_2O$

29. In the following reaction, the product E is



(A) $CH_2OH$–$CHO$ (B) $CHO$–$CO_2H$ (C) $CH_2OH$–$CO_2H$ (D) $CO_2H$–$CO_2H$

**Ans : (C)**

**Hints :** Reaction : $CHO\text{–}CHO \xrightarrow[2.\ H^+]{1.\ NaOH} CH_2OH\text{–}CO_2H$ (Intramolecular Cannizaro's reaction)

30. The amount of electrolytes required to coagulate a given amount f Ag colloidal solution (–ve charge) will be in the order

(A) $NaNO_3 > Al_2(NO_3)_3 > Ba(NO_3)_2$ (B) $Al_2(NO_3)_3 > Ba(NO_3)_2 > NaNO_3$

(C) $Al_2(NO_3)_3 > NaNO_3 > Ba(NO_3)_2$ ( ) $NaNO_3 > Ba(NO_3)_2 > Al_2(NO_3)_3$

**Ans : (D)**

**Hints :** For [AgI] $I^-$ Negatively charged sol, e fec ive i n for coagulation is cation and amount of electrolyte required $\propto \frac{1}{\text{charge content}}$. Also note that $Al(NO_3)_3$ is written as $Al_2(NO_3)_3$ in the questions paper.

31. The value of $\Delta H$ for cooling 2 mole of n deal monoatomic gas from 225°C to 125°C at constant pressure will be given $C_p = \frac{5}{2}R]$

(A) 250 R (B) –500 R (C) 500 R (D) –250 R

**Ans : (B)**

**Hints :** Here, n = 2

$$C_p = \frac{5}{2}R$$

$$\Delta T = 125 - 225 = -100$$

$$\Delta H = nC_p\,\Delta T = 2 \times \frac{5}{2}R \times (-100) = -500\ R$$

32. The quantity of electricity needed to separately electrolyze 1M solution of $ZnSO_4$, $AlCl_3$ and $AgNO_3$ completely is in the ratio of

(A) 2 : 3 : 1 (B) 2 : 1 : 1 (C) 2 : 1 : 3 (D) 2 : 2 : 1

**Ans : (A)**

**Hints :** $Zn^{2+} + 2e^- \rightarrow Zn$

$Al^{2+} + 3e^- \rightarrow Al$

$Ag^{+} + e^{-} \rightarrow Ag$

$\therefore$ Quantity of electricity required = 2 : 3 : 1

33. The emission spectrum of hydrogen discovered first and the region of the electromagnetic spectrum in which it belongs, respectively are

(A) Lyman, ultraviolet (B) Lyman, visible (C) Balmer, ultraviolet (D) Balmer, visible

**Ans : (D)**

**Hints :** Fact

34. As per de Broglie's formula a macroscopic particle of mass 100 gm and moving at a velocity of 100 cm $s^{-1}$ will have a wavelength of

(A) $6.6 \times 10^{-29}$ cm (B) $6.6 \times 10^{-30}$ cm (C) $6.6 \times 10^{-31}$ cm (D) $6.6 \times 10^{-32}$ cm

**Ans : (C)**

**Hints :** m = 100 g, v = 100 cm $s^{-1}$ = 1 $ms^{-1}$

$$\lambda = \frac{h}{mv} = \frac{6.626\times10^{-34}}{0.1\times1} = 6.626 \times 10^{-33} \text{ m} = 6.626 \times 10^{-31} \text{ cm}$$

35. The electronic configuration of Cu is

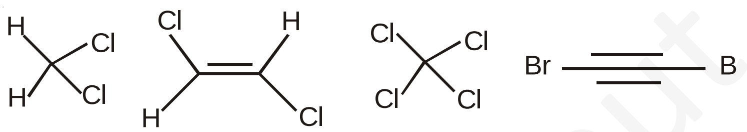
(A) $Ne3s^2 3p^6 3d^9 4s^2$ (B) $Ne3s^2 3p^6 3d^{10} 4s^1$ (C) $Ne3s^2 3p^6 3d^3 4s^2 4p^6$ (D) $Ne3s^2 3p^6 3d^5 4s^2 4p^4$

**Ans : (B)**

**Hints :** Cu : z = 29

[Ne] $3s^2 3p^6 3d^{10} 4s^1$

36. The compound that will have a permanent dipole moment among he following is

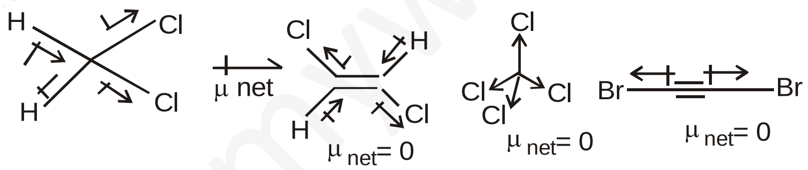



**I** **II** **III** **IV**

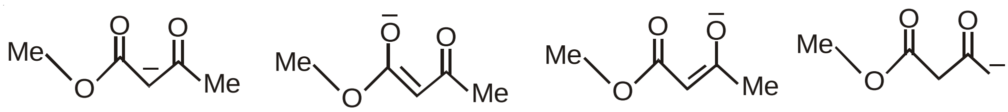
(A) I (B) II (C) III (D) IV

**Ans : (A)**

**Hints :**




37. Among the following structures the one which is not a resonating structure of others is

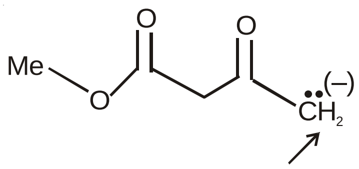



**I** **II** **III** **IV**

(A) I (B) II (C) III (D) IV

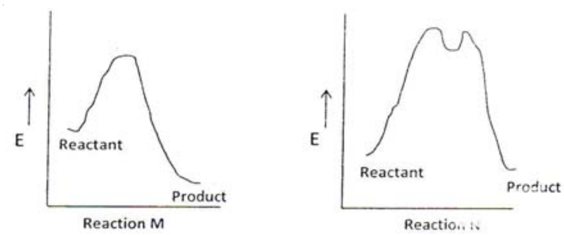
**Ans : (D)**

**Hints :**




A hydrogen is removed from this carbon. But, in resonating structure, position of atoms do not changes.

38. The correct statement regarding the following energy diagrams is




(A) Reaction M is faster and less exothermic than Reaction N

(B) Reaction M is slower and less exothermic than Reaction N

(C) Reaction M is faster and more exothermic than Reaction N

(D) Reaction M is slower and more exothermic than Reaction N

**Ans : (C)**

**Hints :**

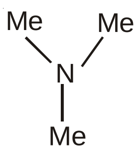
Activation energy ($\Delta E_M < \Delta E_N$)

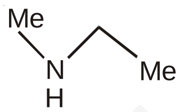
Reaction M is faster than N.

$\Delta H_M$ is more negative than $\Delta H_N$

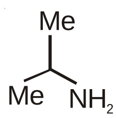
Reaction M is more extothermic than N

39. An amine $C_3H_9N$ reacts with benzene sulfonyl chloride to form a white pr  cipitate which is insoluble in aq. NaOH. The amine is

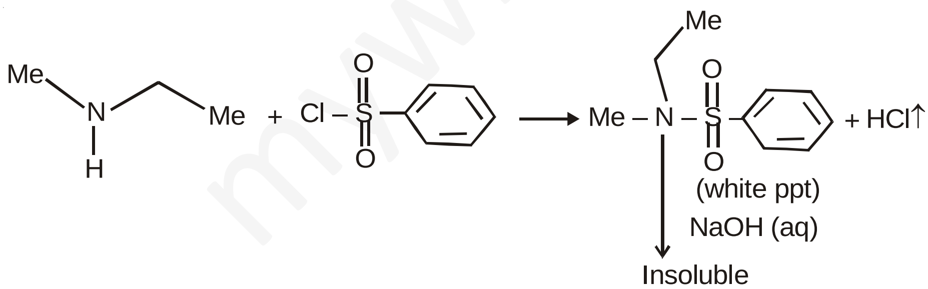
(A) 


(B) 


(C) 


(D) 


**Ans : (B)**

**Hints :** 


40. Among the followings, the one which is not a "greenhouse gas", is

(A) $N_2O$ (B) $CO_2$ (C) $CH_4$ (D) $O_2$

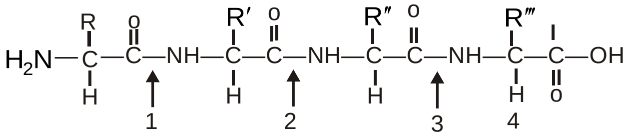
**Ans : (D)**

**Hints :** $O_2$ is not a green house gas

41. The number of amino acids and number of peptide bonds in a linear tetrapeptide (made of different amino acids) are respectively

(A) 4 and 4 (B) 5 and 5 (C) 5 and 4 (D) 4 and 3

**Ans : (D)**

**Hints :** 

No. of amino acids = 4

No. of Peptide bonds = 3

42. The 4th higher homologue of ethane is

(A) Butane (B) Pentane (C) Hexane (D) Heptane

**Ans : (C)**

**Hints :** homologus differ by $CH_2$ unit

$\therefore$ 4th homologue of ethene is $C_6H_{14}$ $\left\{C_2H_6+\left(CH_2\right)_4\right\}$

43. The hydrides of the first elements in groups 15 - 17, namely $NH_3$, $H_2O$ and HF respectively show abnormally high values for melting and boiling points. This is due to

(A) small size of N, O and F

(B) the ability to form extensive intramolecular H-bonding

(C) the ability to form extensive intramolecular H-bonding

(D) effective van der Walls interaction

**Ans : (B)**

**Hints :** $NH_3$, $H_2O$ and HF form extensive intermolecular Hydrogen bonding due to high ionic potential of N, O and F.

44. The two half cell reactions of an electrochemical cell is given as

$Ag^+ + e^- \rightarrow Ag$ ; $E^0_{Ag+/Ag} = -0.3995\ V$

$Fe^{++} \rightarrow Fe^{+++} + e^-$ ; $E^0_{Fe^{+++}/Fe^{++}} = -0.7120\ V$

(A) –0.3125 V (B) 0.3125 V (C) 1.114 V (D) –1.114 V

**Ans : (B)**

$$Ag^+ + e \rightarrow Ag \quad -0.3995\ V\ (cathode)$$

$$Fe^{+2} - e \rightarrow Fe^{+3} \quad -(-0.7120)V\ (An\ \ de)$$

**Hints :**

$$Ag^+ + Fe^{+2} \rightarrow Ag + Fe^{+3} \quad \Delta E = 0.3125\ V$$

$E^o$ cell $= E^o_C - E^o_A$ $E^o$ cell= 0.3125 V

45. In case of heteronuclear diatomics of the type AB, where A is more electronegative than B, bonding molecular orbital resembles the character of A more than that of B. The statement

(A) is false

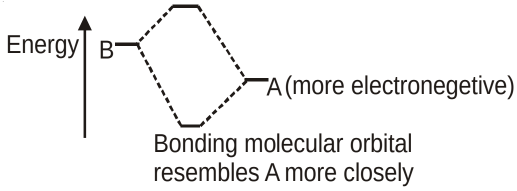
(B) is true

(C) can not be evaluated since data is not sufficient

(D) is true only for certain systems

**Ans : (B)**

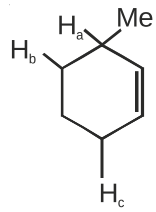
**Hints :**



## CATEGORY - II

**Q.46 to Q.55 carry two marks each, for which only one option is correct. Any wrong answer will lead to deduction of 2/3 mark**

46. The order of decreasing ease of abstraction of Hydrogen atoms in the following molecule is

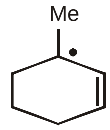



(A) $H_a > H_b > H_c$ (B) $H_a > H_c > H_b$

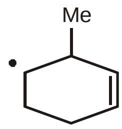
(C) $H_b > H_a > H_c$ (D) $H_c > H_b > H_a$

**Ans : (B)**

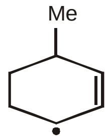
**Hints :** The more stable is the radical formed after H atom abstraction, easier is the abstraction




radical after $H_a$ abstraction (tertiary allyl radical)




radical after $H_b$ abstraction (secondary radical)




radical after $H_c$ abstraction (secondary allyl adica

stability order of free adical is 3 allyl > 2 allyl > 2 a kyl

$\therefore\ H_a > H_c > H_b$

47. The bond angle in $NF_3$(102.3°) is smaller than $NH_3$(107.2°). This is because of

(A) large size of F compared to H (B) large size of N compared to F

(C) opposite polarity of N in the two molecules (D) small size of H compared to N

**Ans : (C)**

**Hints :** In $NF_3$, dipole moment vector point in the direction of F. Thus electron cloud shifts towards F in N–F bond. This reduces bond pair-bond pai r pulsi n in N–F and hence a decrease in bond angle FNF.

48. The compressibility factor (Z) of one mole of a van der Waals gas of negligible 'a' value is

(A) 1 (B) $\frac{bp}{RT}$ (C) $1+\frac{bp}{RT}$ (D) $1-\frac{bp}{RT}$

**Ans : (C)**

**Hints:** <u>Vander Waal's Equation</u>

$\left(P+\frac{a}{V^2}\right)(V-b)=RT$ (for 1 mole of gas) $\Rightarrow$ P(V–b) = RT$\Rightarrow$PV – Pb = RT $\Rightarrow$ PV = RT + Pb$\Rightarrow$ $Z=\frac{PV}{RT}=1+\frac{Pb}{RT}$

Z= Compressibility on neglecting "a".

49. At 25°C, the molar conductance of 0.007 M hydrofluoric acid is 150 mho $cm^2mol^{-1}$ and $\Lambda^\circ_m = 500$ mho $cm^2mol^{-1}$. The value of the dissociation constant of the acid at the gas concentration at 25°C is

(A) $7\times10^{-4}$ M (B) $7\times10^{-5}$ M (C) $9\times10^{-3}$ M (D) $9\times10^{-4}$ M

**Ans : (D)**

**Hints :** $\alpha$(degree of dissociation) = $\frac{150}{500}=0.3$ $\therefore$ $K_a = \frac{C\alpha^2}{1-\alpha}=\frac{0.007\times(0.3)^2}{1-0.3}=9\times10^{-4}M$.

Here, $\alpha$ can't be neglected w.r.t 1 due to large value

50. A piece of wood from an archaeological sample has 5.0 counts min$^{-1}$ per gram of C-14, while a fresh sample of wood has a count of 15.0 min$^{-1}$ gram$^{-1}$. If half life of C-14 is 5770 years, the age of the archaeological sample is

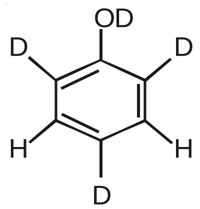
(A) 8,500 years (B) 9,200 years (C) 10,000 years (D) 11,000 years

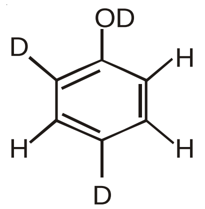
**Ans : (B)**

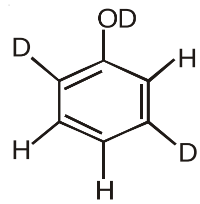
**Hints :** $\frac{0.693}{t_{1/2}} t = 2.303 \log \frac{[\text{Activity of fresh sample}]}{[\text{Activity of fossil}]}$, $\frac{0.693}{5770} t = 2.303 \log \frac{15}{5} \Rightarrow t = \frac{2.303(\log 3)(5770)}{0.693}$ yrs

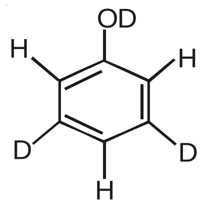
= 9,200 Yrs (approx)

51. When phenol is treated with $D_2SO_4/D_2O$, some of the hydrogens get exchanged. The final product in this exchange reaction is

(A) 

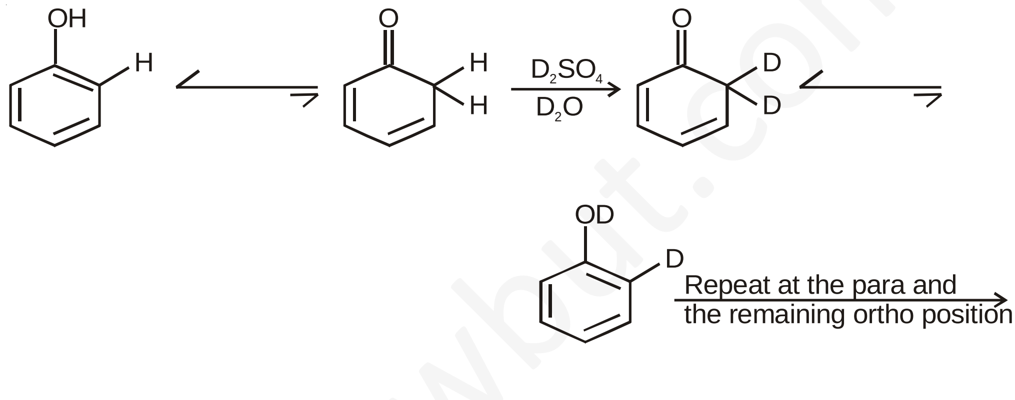
(B) 

(C) 

(D) 

**Ans : (A)**

**Hints :**




52. To observe an elevation of boiling point of 0 05°C, the amount of solute (Mol. Wt. = 100) to be added to 100 g of water ($K_b$=0.5) is

(A) 2 g (B) 0.5 g (C) 1 g (D) 0.75 g

**Ans : (C)**

**Hints :** $\Delta T_b = K_b$ m, 0.05 = .5 × X 0.05= $\frac{0.5x}{100} \times 10$ ; X = 1 g.

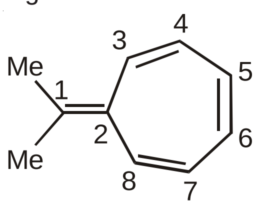
53. The structure of $XeF_6$ is experimentally determined to be distorted octahedron. Its structure according to VSEPR theory is

(A) Octahedron (B) Trigonal bipyramid (C) Pentagonal bipyramid (D) Tetragonal bipyramid

**Ans : (C)**

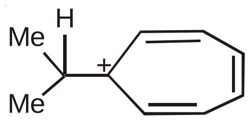
**Hints :** Xe is surrounded by 6 bond pairs and one lone pair. The geometry (geometry of electron pairs) is pentagonal bipyramid.

54. The most likely protonation site in the following molecule is




(A) C-1 (B) C-2 (C) C-3 (D) C-6

**Ans : (A)**

**Hints :**  Aromatic as well as tertiary carbocation

55. The volume of ethyl alcohol (density 1.15 g/cc) that has to be added to prepare 100 cc of 0.5 M ethyl alcohol solution in water is

(A) 1.15 cc (B) 2 cc (C) 2.15 cc (D) 2.30 cc

**Ans : (B)**

**Hints :** Mass of ethyl alcohol before and after the preparation must be equal.

x(volume in cc) × $\frac{1.15g}{mL} = \frac{100 \times 0.5}{1000} \times 46$, x = 2 cc

**CATEGORY - III**

**Q.56 to Q.60 carry two marks each, for which one or more than one options may be correct. Marking of correct options will lead to a maximum mark of two on pro rata basis. There will be no negative marking for these questions. However, any marking of wrong option will lead to award of zero mark against the respective question – irrespective of the number of correct options marked.**

56. Cupric compounds are more stable than their cuprous counterparts in solid state. This is because

(A) the endothermic character of the $2^{nd}$ I P of Cu is not so high

(B) size of $Cu^{2+}$ is less than $Cu^{+}$

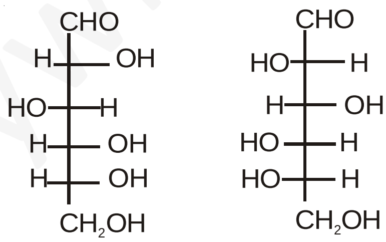
(C) $Cu^{2+}$ has stabler electronic configuration as compared to Cu

(D) the lattice energy released for cupric compounds is much higher than $Cu^{+}$

**Ans : (A, B, D)**

**Hints :** Actually $2^{nd}$ IP of Cu (1958 kJ/mol) is not very high as compared to 1st IP (745 kJ/mol). In addition the gain in lattice energy due to +2 state and small size of Cu favour the divalent state in the solid.

57. Among the following statements about the molecules X and Y, the one (s) which is (are) correct is (are)




X Y

(A) X and Y are diastereomers (B) X and Y are enantiomers

(C) X and Y are both aldohexoses (D) X is a D-sugar and Y is an L-sugar

**Ans : (B, C, D)**

**Hints :** 'X' and 'Y' are mirror images of each other. They are aldohexoses too. In 'X', –OH of the asymmetric 'C' farthest from –CHO is on the right, so it is 'D'-Sugar. 'Y' , on the other hand, has –OH on the left. Thus it is a L-sugar.

58. For a spontaneous process, the correct statement(s) is (are)

(A) $(\Delta G_{system})_{T,P} > 0$ (B) $(\Delta S_{system}) + (\Delta S_{surroundings}) > 0$

(C) $(\Delta G_{system})_{T,P} < 0$ (D) $(\Delta U_{system})_{T,V} > 0$

**Ans : (B, C)**

**Hints :** Spontaneity of of the process can be expressed either by taking entropy changes of system and surrounding together or by considering free energy change of the system alone at constant temperature and pressure. The known criteria are : $(\Delta G_{sys})_{T,P} < 0$ and $(\Delta S_{sys}) + (\Delta S_{sur}) > 0$

59. The formal potential of $Fe^{3+}/Fe^{2+}$ in a sulphuric acid and phosphoric acid mixture (E°=+0.61V) is much lower than the standard potential (E°=+0.77V). This is due to

(A) formation of the species $[FeHPO_4]^+$

(B) lowering of potential upon complexation

(C) formation of the species $[FeSO_4]^+$

(D) high acidity of the medium

**Ans : (A, B, D)**

**Hints :** Formation of complex by $Fe^{3+}$ reduces its concentration. Thereby lowers the formal reduction potential.

60. Two gases X (Mol. Wt. $M_x$) and Y(Mol. Wt. $M_Y$; $M_Y > M_x$) are at the same temperature T in two different containers. Their root mean square velocities are $C_X$ and $C_Y$ respectively. If the average kinetic energies per molecule of two gases X and Y are $E_x$ and $E_Y$ respectively, then which of the following relation (s) is (are) true?

(A) $E_x > E_Y$

(B) $C_x > C_Y$

(C) $E_x = E_Y = \frac{3}{2}RT$

(D) $E_X = E_Y = \frac{3}{2}k_BT$

**Ans : (B, D)**

**Hints :** For same temperature, higher the molar mass, lower is the rms velocity.KE of individual molecules is expressed in terms of $K_B$ not R

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