List of 50 Important Chemistry Questions and Answer Download in PDF:

The List of 50 important Chemistry Questions and Answers is given here for download in PDF. Candidates those who are preparing for SSC CGL and all other competitive exams can use these questions.

1). Solutions are classified into aqueous and non-aqueous solutions, based on_____.

a) Nature of solute particles
b) Nature of solvent
c) Size of the particles
d) Thickness of solvent

Answer is: b)

2). The solvent used to prepare aqueous solutions is_____.

a) Water
b) benzene
c) kerosene
d) petrol

Answer is: a)

3). A true solution does not show Tyndall effect, because of the_____.

a) Nature of solvent
b) Amount of solute
c) Size of the particles
d) Nature of solute

Answer is: c)

4). Tyndall effect is exhibited by_____.

a) True solutions

b) Suspensions

c) Colloidal solutions

d) Crystals

Answer is: c)

5). Tyndall effect is producted by_____.

a) True solutions of light

b) Scattering of light

c) Refraction of light

d) Movement of particles

Answer is: b)

6). The particle size in a colloidal solution is _____.

a) 1 Å – 10 Å

b) 10 Å - 2000 Å

c) More than 2000 Å

d) Less than 1 Å

Answer is: b)

7). The particle size in a suspension is_____.

a) 1 Å – 10 Å

b) 10 Å - 2000 Å

c) More than 2000 Å

d) Less than 1 Å

Answer is: c)

8). A solution which has more of solute, at a given temperature than that of saturated solution is called a_____.

a) Super saturated solution
b) Unsaturated solution
c) Colloidal solution
d) suspension

Answer is: a)

9). Chalk powder in water is an example of_____.

a) Saturated solution
b) Unsaturated solution
c) suspension
d) Colloidal solution

Answer is: c)

10). The particle size of the solute in true solution is_____.

a) 1 Å – 10 Å
b) 10 Å - 100 Å
c) 100 Å - 1000 Å
d) More than1000 Å

Answer is: a)

11).Milk is a _____.

a) True solution

b) Colloidal solution

c) suspension

d) saturated solution

Answer is: b)

12).Nitrogen in soil is an example for_____.

a) True solution

b) saturated

c) super saturated

d) unsaturated

Answer is: b)

13).Fog is a solution of _____.

a) Liquid in gas

b) Gas in liquid

c) Solid in gas

d) Gas in gas

Answer is: a)

14).Soda water is a solution of _____.

a) Liquid in gas

b) Gas in liquid

c) Solid in gas

d) Gas in gas

Answer is: b)

15).Blood is an example of _____.

a) True solution
b) Colloidal solution
c) Saturated solution
d) Suspension

Answer is: b)

16).The dispersed phase in a colloidal solution is_____.

a) Solute
b) Solution
c) Suspension
d) Mixture

Answer is: a)

17).Sugar and Salt solutions are_____.

a) Heterogeneous mixtures
b) True solutions
c) Colloidal solutions
d) Suspensions

Answer is: b)

18).Brownian movement explains the _____ property of colloidal solutions.

a) optical
b) electrical
c) kinetic

d) mechanical

Answer is: c)

19).In aqueous solutions, the solvent used is_____.

a) benzene
b) ether
c) alcohol
d) water

Answer is: d)

20).The solution in which saturation is not achieved is called _____.

a) Super saturated
b) Unsaturated
c) Saturated
d) Suspended

Answer is: b)

21).Cheese is a colloidal solution of _____.

a) Solid in solid
b) Liquid in solid
c) Solid in liquid
d) Gas in solid

Answer is: b)

22).Cork is a colloid of _____.

a) Solid in solid
b) Liquid in solid
c) Solid in liquid

d) Gas in solid

Answer is: d)

23).Smoke is a colloid of _____.

a) Solid in solid

b) Liquid in solid

c) Solid in liquid

d) Solid in Gas

Answer is: d)

24).The saturation temperature for 20.7g of $CuSO_4$ soluble in water is_____.

a) 10^0C

b) 100^0C

c) 20^0C

d) 30^0C

Answer is: c)

25).The solubility level of an aqueous solution of NaCl at 25^0C is _____.

a) 20g

b) 36g

c) 95g

d) 8g

Answer is: b)

26).The increase in the solubility of Sodium halides, in water at 25^0C is_____/

a) NaCl > NaBr > NaI

b) NaBr > NaI > NaCl

c) NaI > NaBr > NaCl

d) NaCl = NaBr > NaI

Answer is: c)

27).Solubility of CaO in water is a _____.

a) Chermic

b) endothermic

c) exothermic

d) hypothermic

Answer is: c)

28).According to Henry's Law, in gases, an increase in pressure increase_____.

a) Solubility

b) saturation

c) volume

d) viscosity

Answer is: a)

29).Deep sea divers use mixture of _____.

a) Helium - Oxygen

b) Nitrogen - Oxygen

c) Hydrogen - Nitrogen

d) Helium - Nitrogen

Answer is: a)

30).The continuous random motion of colloidal particles is called _____.

a) Brownian movement

b) Zig zag movement

c) Continuous movement

d) Tyndall effect

Answer is: a)

31).On increasing the temperature, the solubility of the solute in the solvent_____.

a) Increase

b) Decrease

c) Change

d) Does not change

Answer is: a)

32).Which law relates solubility of solvents with pressure?

a) Hess' law

b) Henry's law

c) Charles' Law

d) Boyle's law

Answer is: b)

33).When sunlight passes through the window of your house, the dust particles scatter the light making the path of the light visible. This phenomenon is called as_____.

a) Brownian motion

b) Tyndall effect

c) Raman effect

d) Uniform motion

Answer is: b)

34).The Greek term 'atomos' means _____.

a) divisible

b) indivisible

c) macro molecule

d) soft sphere

Answer is: b)

35).Isotopes are the atoms of same element, with same atomic number. But with different _____.

a) Atomic number
b) Mass number
c) Number of electrons
d) Chemical nature

Answer is: b)

36).$_6C^{12}$ and $_6C^{14}$ are _____.

a) Isotopes
b) Isobars
c) Isomers
d) Molecules

Answer is: a)

37).Atoms of different elements possessing in the same atomic mass are called _____.

a) Isotopes
b) Isobars
c) Isomers
d) Molecules

Answer is: c)

38).Atoms of different elements with same number of neutrons.

a) Isotopes

b) Isomers

c) Isobars

d) Isotones

Answer is: d)

39).Atomicity of oxygen in ozone molecule is_____.

a) 1

b) 2

c) 3

d) 4

Answer is: c)

40).Atomicity of primary gases is_____.

a) 1

b) 2

c) 3

d) 4

Answer is: b)

41).In the Beginning of the 20th century, Matter Wave concept was introduced by_____.

a) Broglie

b) Avogadro

c) Heisenberg

d) Einstein

Answer is: a)

42).The Principle of Uncertainty was introduced by_____.

a) Broglie
b) Avogadro
c) Heisenberg
d) Einstein

Answer is: c)

43).$_{18}Ar^{40}$ and $_{20}Ca^{40}$ are considered as_____.

a) Isotopes
b) Isomers
c) Isobars
d) Isotones

Answer is: a)

44).The compound which does not show simple ratio of atoms, is _____.

a) Benzene
b) Acetylene
c) Hydrogen
d) Sucrose

Answer is: d)

45).Avogadro's hypothesis relates volume of gases and _____.

a) mass
b) temperature

c) pressure

d) number of molecules

Answer is: d)

46).Atomicity of an element is _____.

a) Valency of an element

b) Atomic mass

c) Number of atoms in one molecule of an element

d) Isotope of an element

Answer is: c)

47).Atomicity is given by_____.

a) Mass/molecular mass

b) Mass of the element

c) Molecular mass X atomic mass

d) Molecular mass / atomic mass

Answer is: d)

48).The atoms of $_{6}C^{13}$ and $_{7}N^{14}$ are considered as _____.

a) Isotopes

b) Isomers

c) Isobars

d) Isotones

Answer is: d)

49).Isotones are the atoms of different elements having_____.

a) Same mass number

b) Same atomic number

c) Same number of neutrons

d) Same number of electrons

Answer is: c)

50).Atomicity of Phosphorous is _____.

a) 2

b) 3

c) 4

d) 5

Answer is: c)